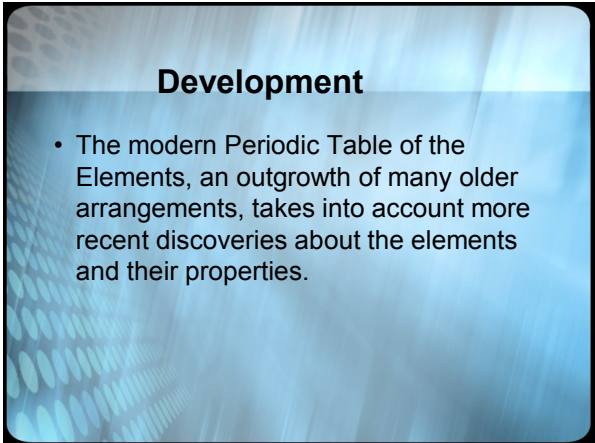


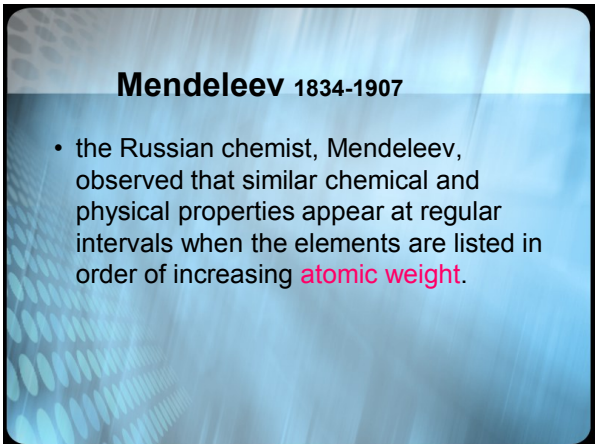
The Periodic Table

History and Organization



Development

- The modern Periodic Table of the Elements, an outgrowth of many older arrangements, takes into account more recent discoveries about the elements and their properties.



Mendeleev 1834-1907

- the Russian chemist, Mendeleev, observed that similar chemical and physical properties appear at regular intervals when the elements are listed in order of increasing **atomic weight**.

Moseley 1885-1915

- X-ray studies done by Moseley, an English Physicist, showed that properties of the elements are a function of the **atomic number**, not of the atomic weight.

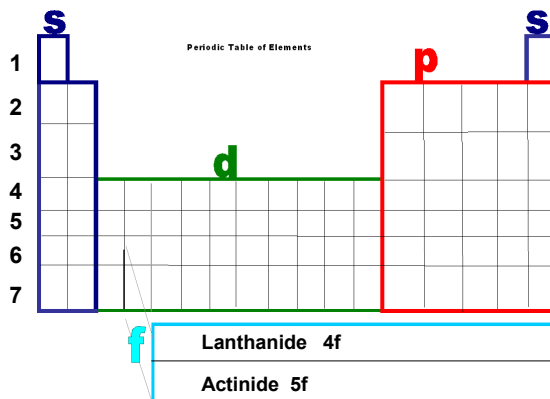


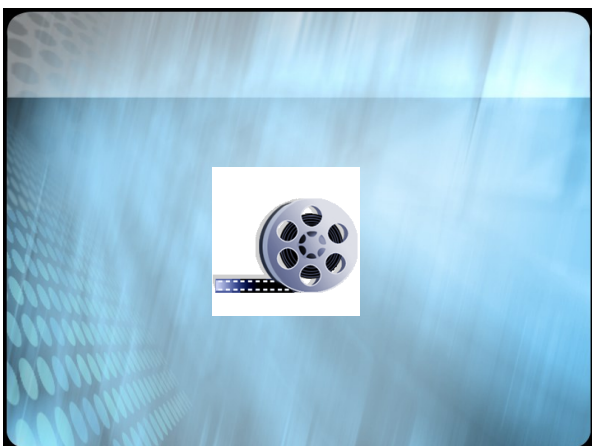
Periodic Law

- The modern Periodic Law states :
- *The properties of the elements are periodic functions of their atomic numbers.*

PERIODS

- The number of each period shows the principle energy level



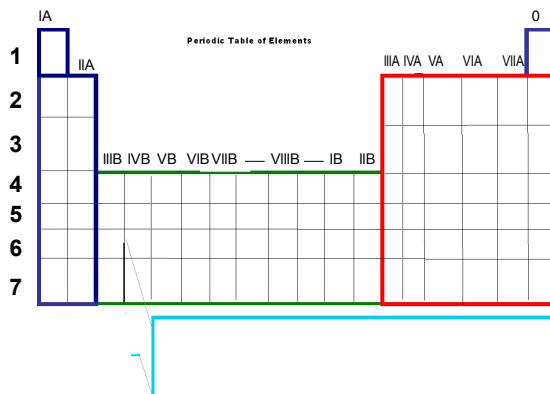


GROUPS

- The vertical columns of the Periodic Table are called **groups** or **families**.

Group Designations

- The traditional designation for each group has been a combination of a Roman numeral and the letter A or B, such as IA or IIB.
- These designations are often used along side of a new form which numbers the groups from 1 to 18.



Groups

